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Review Objectives • State a rule that  
usually tells how many electrons are  
shared to form a covalent bond •  
Describe how electron dot formulas are  
used • Predict when two atoms are likely  
to be joined by a double or a triple  
covalent bond • Distinguish

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Experimental evidence, however,  
indicates ...



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between atoms. a neutral group of atoms joined together by covalent bonds. a molecule consisting of two atoms. the mnemonic for remembering which non-metal elements form dia....

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Bonding. Covalent Bond. Molecule.  
Diatomic Molecule. Monatomic Molecule.  
A chemical bond that involves sharing a  
pair of electrons betw.... The smallest  
particle of a substance that retains the  
chemical.... A molecule that consists of  
two atoms of the same element that....

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Lecture Presentation

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Exercise 2 Lewis Structure of a  
Compound. Solve. Nitrogen must share a  
pair of electrons with three fluorine

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atoms to complete its octet. Thus, the binary compound these two elements form must be  $\text{NF}_3$ . Given the Lewis symbols for nitrogen and fluorine in Table 8.1, predict the formula of the stable binary compound (a compound composed of two elements) formed when nitrogen reacts with ...

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